

Short Communication

## Determination of the Real Surface Area of a Screen-Printed Electrode by Chronocoulometry

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Received: 27 April 2012 / Accepted: 25 May 2012 / Published: 1 July 2012

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A simple and effective method applying chronocoulometry and cyclic voltammetry to the determination of the real surface area of a screen-printed electrode, for the first time, is reported. The method is based on the Anson plot and the standard addition method. This method could greatly facilitate determination of the real surface area of commercial screen-printed electrodes, which still remains underdeveloped despite the critical importance of quality control in this industry.

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**Keywords:** Screen-Printed Electrodes, Real/Active Surface Area, Chronocoulometry, Anson Plot, Standard Addition Method

### 1. INTRODUCTION

The actual surface area of an electrode where electron exchange takes place is termed the real or active electrochemical surface area. It is different from the geometrical surface area, which is simply the sum of all the physical areas that cover the surface of the electrode. The ratio of the electrochemical surface to the geometrical surface is represented by the roughness factor ( $\rho$ ) as a coefficient [1]. In general, the active area of an electrode exceeds the geometrical surface area creating a higher number of chemically reactive sites on comparatively small electrodes. Most of the kinetic parameters of the electrode reaction as well as the electrical double layer properties need to be referred to the unit area. Therefore, the determination of the real surface area of the electrodes plays a crucial role in the calculation of various parameters characterising electrochemical processes [2] and facilitates quality control of mass-produced electrodes.

Various electrochemical methods that determine the real surface area of conventional solid electrode (*e.g.* carbon electrodes and metal electrodes) have been reported [3]. Furthermore, Trasatti and Petri described many methods for the determination of real electrochemical surface area of liquid electrodes [4]. By contrast, however, the determination of the real electrochemical surface area of screen-printed electrodes still remains underdeveloped, which is mainly due to the inherent physical properties of screen-printed electrodes. Thus, some electrochemical methods require good surface condition and are therefore inapplicable to the determination of the real surface area of screen-printed electrodes [5]. In addition, there are some other hindrances such as the complexity and non-uniformity of the materials and production procedures for screen-printed electrodes. To the best of our knowledge, no effective method has previously been reported for the determination of the real electrochemical surface area of screen-printed electrodes.

## 2. EXPERIMENTAL

### 2.1 Materials (Apparatus & Reagents)

TMPD (*N,N,N',N'*-tetramethyl-*p*-phenylenediamine) was purchased from Acros-Organics (Geel, Belgium). 2-(1-Methoxy)propyl acetate 97%, used as diluter for preparing stock solutions of TMPD, was purchased from Acros-Organics (Geel, Belgium). All reagents used to make PBS (phosphate buffer saline) buffers were purchase from Sigma-Aldrich (Dorset, UK). Deionised water was used in all solution preparations.

The commercial screen-printed electrodes applied in this study were supplied by Pelikan Technologies GmbH (Muenster, Germany). The screen-printed electrodes were fabricated on polyvinyl chloride (PVC) sheets using a screen-printer. Each electrode comprised a three electrode system: a printed Ag/AgCl reference electrode a Ag/AgCl counter electrode and a carbon electrode acting as the working electrode.

### 2.2 Methods

All of the cyclic voltammetric and chronocoulometric experiments were conducted using a  $\mu$ -Autolab type III potentiostat (Eco Chemie, Netherlands). The final data included in this communication were the mean of the measurements in triplicate, except the chronocoulometric experiment where 5 replicates had been run. All solutions were degassed by bubbling pure nitrogen before use.

The diffusion coefficient of the mediator was calculated using cyclic voltammetry. After adding 0.7  $\mu$ L of 10 mM pH 7.4 PBS on a screen-printed electrode, cyclic voltammograms were run with three different scan rates (10, 25 and 50 mV/sec). For the determination of unknown mediator concentration in a screen-printed electrode, six TMPD solutions with different concentrations (0, 2, 4, 6, 8 and 10 mM) in PBS were first prepared as the standard additions. The TMPD mediator was initially dissolved in the diluter, 2-(1-Methoxy)propyl acetate 97% and it was further diluted with

10mM PBS buffer (pH 7.4), in order to achieve a completely dissolved and homogenous solution. An aliquot (0.7  $\mu\text{L}$ ) of each concentration was then added on the electrode and cyclic voltammograms were run with a scan rate of 50mV/sec.

The chronocoulometric experiments were conducted by applying two potentials. The pulse width was 2 sec. The first potential (-0.25 V) was low so that no electrochemical reduction of TMPD occurred, and the second potential (0.25 V) applied was high enough to reduce all the TMPD. After adding 0.7  $\mu\text{L}$  of 10mM pH 7.4 PBS on the electrode, the total charge passed versus time was recorded for 4 seconds.

SEM images of the working electrode paste were obtained using a ISI ABT-55 low-vacuum scanning electron microscope. The working electrode paste provided by Pelikan was spread to an even film on a microscope glass slide by using a plastic spatula. After being dried for 2 hours in oven at 65  $^{\circ}\text{C}$ , the slide was tailored and placed on a cylinder holder for the observation.

### 3. RESULTS & DISCUSSION

In this communication, we describe a simple and effective methodology utilising chronocoulometry for the determination of the real electrochemical surface area of a screen-printed electrode for the first time. This study focuses on the screen-printed carbon electrode, and the method explored is based on the Anson plot and the standard addition method.

Chronocoulometry (CC) is one of the classical electrochemical techniques used in electroanalytical chemistry. Briefly speaking, it is the measurement of charge (coulombs) as a function of time. In other words, a fast-rising potential pulse is applied to the working electrode and the electrical charge passing through this electrode is measured as a function of time [6]. The technique of chronocoulometry has been successfully applied to the surface area measurement of conventional solid electrode, kinetics of both heterogeneous electron transfer reactions and chemical reactions coupled to electron transfer [7].

The analysis of the chronocoulometric data is based on the Anson equation (Eq.1) [8] which defines the charge-time dependence for linear diffusion control:

$$Q = 2nFACD^{1/2}\pi^{1/2}t^{1/2} \quad (1)$$

Herein, Q represents the charge (coulombs). n and A represent the number of electrons transferred and the real electrochemical surface area of the electrode ( $\text{cm}^2$ ), respectively. F is the Faraday's constant (96,485 coulombs/mole). The concentration of the mediator is denoted by C where the diffusion coefficient of the mediator ( $\text{cm}^2/\text{sec}$ ) and time (sec) are denoted by D and t, respectively.

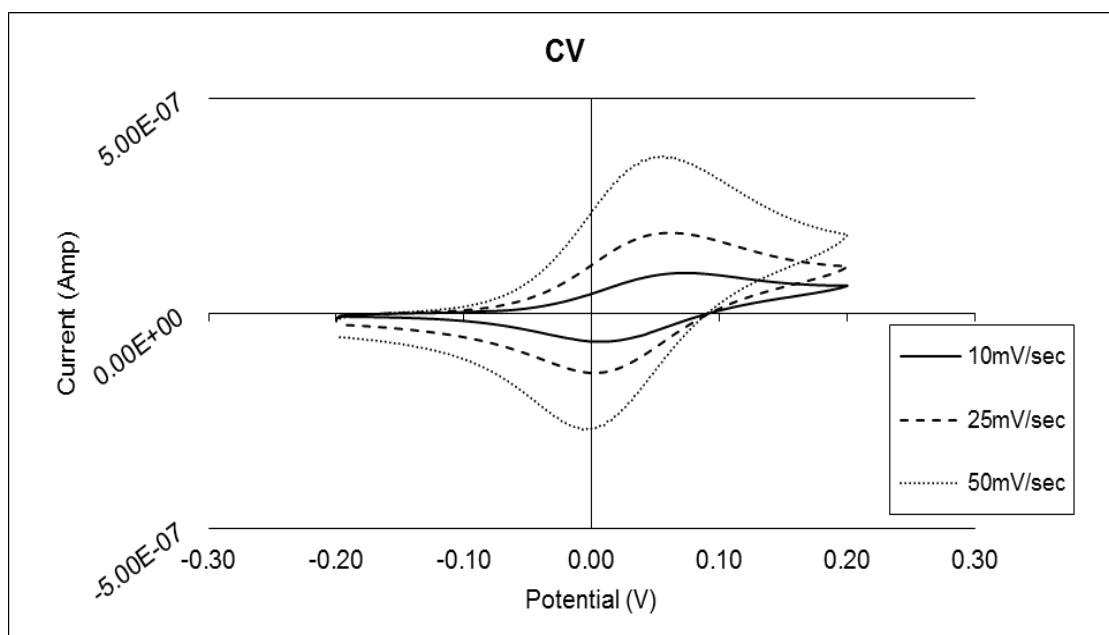
There is a linear relationship between the charge (Q) and the square root time ( $t^{1/2}$ ). A plot of Q vs.  $t^{1/2}$  is often referred as the Anson plot [7] where the slope (a) can be determined by the following equation (Eq.2):

$$a = 2nAFCD^{1/2}/\pi^{1/2} \quad (2)$$

Therefore, the real electrochemical surface area of the electrode (A) can be calculated from the slope in Eq.3, provided the other four parameters (n, F, C and D) are known.

$$A = a/(2nFCD^{1/2}/\pi^{1/2}) \quad (3)$$

Apart from the number of electrons transferred (n) and the Faraday's constant (F) which are already well known, the calculation of the diffusion coefficient (D) of the mediator was achieved by using cyclic voltammetry. As can be seen in **Fig. 1**, after adding 0.7 $\mu$ L phosphate buffer saline (PBS) on a screen-printed electrode, cyclic voltammograms were obtained with three different scan rates and the resulting currents of the oxidation peaks were plotted against the square root of the scan rates. The slope obtained is equal to the diffusion coefficient of the mediator. Thus, the diffusion coefficient of the mediator is 5.98E-08 cm<sup>2</sup>sec<sup>-1</sup> (the coefficient of variation is 15%, while the confidence interval is 1.04%).



**Figure 1.** Cyclic voltammograms of a commercial screen-printed electrode at three different scan rates

The precise concentration of mediator in a typical commercially produced screen-printed electrode is, however, unknown. It is difficult to measure the concentration of the mediator directly since the working electrode paste is a complex matrix containing not only the mediator (in this case TMPD, *N,N,N',N'*-tetramethyl-p-phenylenediamine), but also carbon paste, enzyme (GOx, glucose oxidase), and a series of dried buffer components. Moreover, it is impractical to either separate or remove the mediator from the matrix, since the physical property and structure of the matrix will be changed leading to unpredictable results.

The method of standard addition is widely applied to determine the analyte concentration in a complex matrix if the analytical signal is proportional to the analyte concentration [9]. It can be presented as Eq.4 if the sample response is known to vary linearly with sample concentration [10]

$$C_u = r_u C_s / (r_{u+s} - r_u) \quad (4)$$

Where:

$C_u$ : unknown concentration

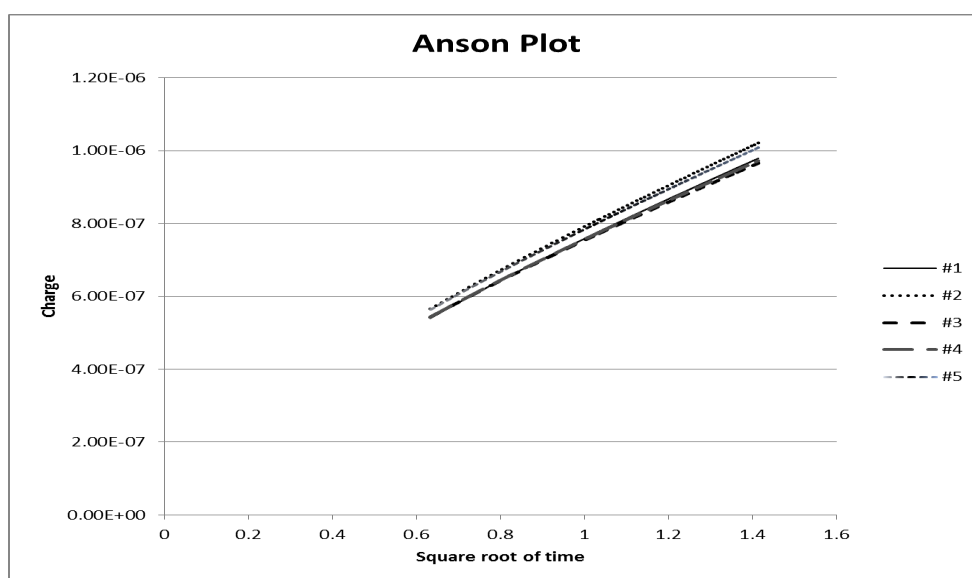
$r_u$ : response of the unknown (analytical signal of the unknown)

$r_{u+s}$ : response of the unknown plus the standard (analytical signal of the unknown plus the standard)

In practice, several aliquots of standard TMPD solution were added to the system. A plot was then created showing the response against the concentration of each addition, and the straight line obtained was then extrapolated back to a zero response. The negative intercept on the concentration axis represents the unknown concentration, which could be calculated by dividing the positive intercept on the response axis by the slope.

Since in our case it is well known that the electrical current of an electrochemical process is linearly proportional to the concentration of the analyte used [11], by using the standard addition method (Eq.4), the unknown concentration ( $C_u$ ) of the mediator TMPD was obtained successfully. The average concentration of the mediator TMPD in the commercial screen-printed electrode was  $3.03 \text{ M}$  ( $3.03 \times 10^{-6} \text{ mol} \cdot \text{cm}^{-3}$ ) (the coefficient of variation is 16%, while the confidence interval is 0.55%).

Consequently, by using the Anson plot (Fig. 2) generated in the following chronocoulometric measurement and the Anson Equation (Eq.3) introduced previously, the real electrochemical surface area of the commercial screen-printed electrode was estimated as  $0.80 \text{ mm}^2$  (the coefficient of variation is 4.25%, while the confidence interval is 0.03%) (Tab. 1).



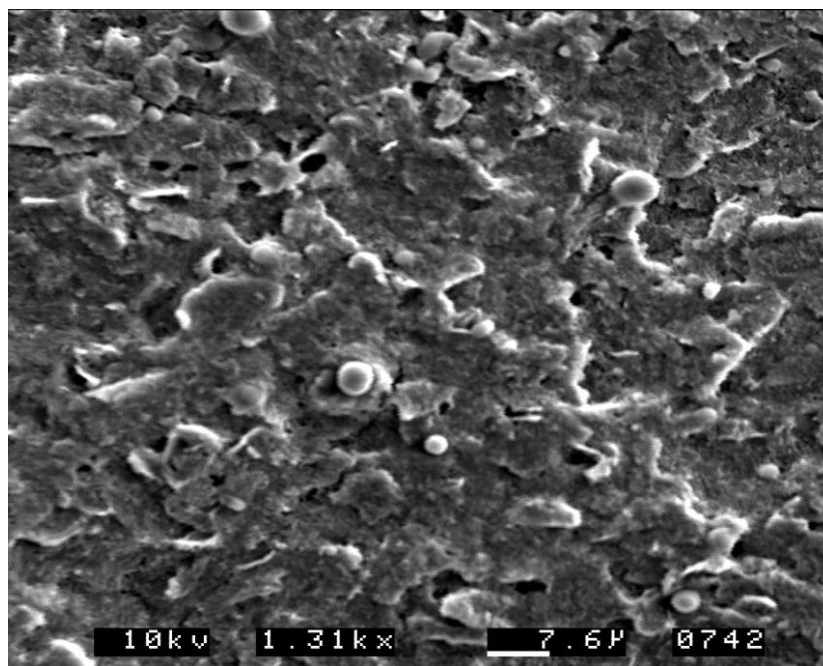
**Figure 2.** Anson plots of a commercial screen-printed electrode (five replicates)

**Table 1.** Real electrochemical surface area of the screen-printed electrodes

Replicates	slope	C ( $\mu\text{M}$ )	C ( $\text{mol}/\text{cm}^3$ )	D( $\text{cm}^2/\text{sec}$ )	A ( $\text{mm}^2$ )	
1	6.29E-07	3027.5	3.03E-06	5.98E-08	0.79	
2	6.24E-07	3027.5	3.03E-06	5.98E-08	0.77	
3	6.88E-07	3027.5	3.03E-06	5.98E-08	0.85	
4	6.44E-07	3027.5	3.03E-06	5.98E-08	0.79	
5	6.65E-07	3027.5	3.03E-06	5.98E-08	0.82	
					0.80	Average
					0.03	STDEV
					4.25	Coefficient of Variation %
					0.03	Confidence Interval (95%)

As can be seen in Tab. 1, the standard deviation of the real electrochemical surface area is quite small. In other words, a batch of screen-printed electrodes showed relatively consistent results for the real electrochemical surface area determination using the above method. As a result, it is suggested that the method could be considered reliable, since the reproducibility between replicate determinations for biosensors should be at least  $\pm$  (5%-10%) [11].

Furthermore, as expected, the real electrochemical surface area of the screen-printed electrode is much larger than the geometrical electrode surface area, which is  $0.25 \text{ mm}^2$ . The roughness factor [12] is 2.72 (ratio of electrochemical surface to the geometrical surface). This roughness factor could be attributed to the morphology of the working electrode paste. The rough surface of the working electrode paste was confirmed by scanning electron microscopy (SEM) in Fig. 3.

**Figure 3.** SEM image of the working electrode paste (magnification at 1.31kx)

#### 4. CONCLUSIONS

In conclusion, we have first demonstrated a simple and effective method using chronocoulometry and cyclic voltammetry for the determination of the real surface area of a screen-printed electrode. The resulting high roughness factor ( $\rho$ ) of the working electrode paste was in agreement with the image of SEM. This method should facilitate determination of the real surface area of commercially produced screen-printed electrodes thus improving quality control in this critically important industry, which serves the needs of people with diabetes.

#### ACKNOWLEDGEMENT

Financial support from Pelikan Technologies, Inc. is gratefully acknowledged.

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