

Short Communication

Phase Transitions Observations During the Direct Electrolysis of Ilmenite to Ferrotitanium in $\text{CaCl}_2\text{-NaCl}$ Melt

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The electrochemical reduction method is successfully used to prepare porous ferrotitanium alloys from ilmenite by using equal-molar $\text{CaCl}_2\text{-NaCl}$ eutectic as electrolyte, molybdenum rod as cathode and graphite as anode at 973K with cell voltage of 3.0V under inert atmosphere. The intermediate products formed in the electrochemical reduction process are analyzed by cyclic voltammetry and constant cell voltage electrolysis in the two electrodes system together with XRD, EDS analysis and SEM observation. It is demonstrated that the electrochemical reduction of ilmenite is a stepwise process since the intermediates CaTiO_3 , Fe_2Ti and Ti are observed in the products partially reduced. In the electrochemical reduction process, the reduction of FeTiO_3 firstly gives rise to the formation of Fe and CaTiO_3 mixtures, which are formed at the interface between the ilmenite pellet and the electrolyte and are agglomerated on the surface of the samples. The intermediate CaTiO_3 will be further reduced in the presence of solid Fe to form ferrotitanium alloys as well as directly deoxidized to titanium metal, which is then combined with Fe_2Ti to form ferrotitanium. This reduction path is in good agreement with the cyclic voltammetry analysis. Experimental results also show that the reduced metallic grains are interconnected on the surface of the unreduced CaTiO_3 particles, indicating that the metallic matrix helps the electron transfer into the oxide inside.

Keywords: Ferrotitanium, ilmenite, electrolysis, cyclic voltammetry.

1. INTRODUCTION

Ferrotitanium alloy, well known as the useful hydrogen-storage alloy and deoxidizer, has been applied commercially in the hydrogen storage and steel-making industries. Traditionally [1-5], ferrotitanium is prepared by the carbothermic and aluminothermic reduction of oxide ores and

concentrates such as ilmenite ores or by re-melting the mixtures of iron and titanium at a proper ratio. However, these methods have many disadvantages such high energy consumption, low quality of products and so on. For example, carbothermic process should be operated in an electric arc furnace at extremely high temperatures, which leads to not only an extensive energy consumption but also large amount of CO₂ emissions. The aluminothermic reduction and re-melting method, however, are regrettably high energy consumption also. Thus, developing both environmentally friendly and energy saving methods seem to be urgent for the production of ferrotitanium alloy.

Over years, many efforts have been focused on the electrometallurgical techniques to produce and purify metals [6-9]. Of the existing electro-refining techniques, the FFC process has been studied worldwide and applied in the successful production of various metals and alloys, such as Nb, Ti, Tb, Zr, TbFe₂, FeTi, CeCo₅, TiMo and ZrSi alloys [7-35]. The mechanism of this process can be summarized as follows: first, experiments are carried out in fused alkaline chloride in the two electrodes system with the metallic oxide pellets as cathode and graphite as anode; second, the oxide pellets are deoxidized at the cathode and the released oxygen ions are diffused in the molten salt and then removed as O₂, CO or CO₂ gas at the anode. Recently, the electrochemical techniques have been applied in the electrochemical reduction analysis and provide reliable evidences particular in the mechanism and kinetics properties of electrode reactions. The classical electrochemical system is composed of three electrodes [11-22]: the molybdenum metallic cavity electrode (MCE) as the working electrode, the graphite crucible as the counter electrode and the Ag/AgCl [36] or Ag wire as the reference electrode. Other electrodes can be also found in literatures [8,37].

The FFC process provides a promising technique for the extraction of FeTi alloys from their oxides. Researchers have successfully prepared FeTi alloys by this method with the pretreated ferrotitanium oxide processor as cathode, graphite as anode and CaCl₂ as electrolyte [25,38-40]. Similar results can be also tracked by the SOM process [41]. The electrochemical reduction mechanism, however, reaches few attentions, especially the intermediate phase transitions as well as the electrochemistry properties could be hardly found in the present literatures. Moreover, it has been found that the intermediate CaTiO₃ was always formed during the reduction. The reducing path of CaTiO₃ is needed more detailed researches to clarify the reduction mechanism.

In this communication, experiments are carried out to study the electrochemical behavior and the phase transitions in the electrochemical reduction of ilmenite by cyclic voltammetry and constant cell voltage electrolysis. Moreover, to study the electrochemical reduction process of ilmenite could further enrich and consummate the reduction mechanism and contribute to the development of the electrochemical extraction of metallic oxides technique in molten salts.

2. EXPERIMENTAL

2.1 Preparation of CaCl₂-NaCl Eutectic Electrolyte

The as-received analytical grade anhydrous CaCl₂ (73g, purity >97 mass%) and NaCl (32g, purity >98 mass%), mixed at equal molar ratio, were added into a Al₂O₃ crucible (I.D. 57mm, O.D.

63mm, and 75mm in depth) and then transferred into a vertical quartz tube reactor under argon (analytical grade, >99.999%). The mixture was heated to 973K that was higher than the eutectic point of 777K. Pre-electrolysis was carried out at 2.8V for 20h between the graphite anode (99.9% in purity, dia. 6mm, 80mm in length) and molybdenum plate (10mm×10mm) cathode in order to remove redox-active impurities remained in the molten salt.

2.2 Cyclic voltammetry analysis

Cyclic voltammetry analysis was recorded by using a computer controlled CHI760D Electrochemical System (Shanghai Chenhua, China). The three electrodes system containing the metallic cavity electrode (MEC) [42] filled with/without FeTiO_3 powder as the working electrode, a graphite stick (99.9% in purity, dia. 2mm, 100mm in length) as the counter electrode, and a Ag wire (dia. 0.5mm, 50mm in length) as the quasi-reference electrode. Generally, the metallic cavity electrode, named widely as the MCE, has been employed comprehensively in the cyclic voltammetry analysis [43-48] due to at least two advantages. One is the decrease of the iR drop and the other is the consequently enhancement of the CV measurements sensitivity [6]. In the present study, the MCE was fabricated by high-speed mechanical drilling a $\sim 0.35\text{mm}$ circular hole at the end of molybdenum foil (0.5mm thickness, $\sim 0.7\text{mm}$ width and $\sim 100\text{mm}$ in length). The foil was then polished on the metallographic sandpaper, boiled in concentrated NaOH for another 10 minutes, washed in distilled water and dried in air. Then, FeTiO_3 powder (particle size ~ 50 micrometers) was manually filled into the MCE hole by repeatedly finger-pressing the powder on both sides of the foil, and wiped off carefully any powder left on the surface of the foil. The corresponding optical photographs for MCE with/without FeTiO_3 powder, were shown in Fig. 1.

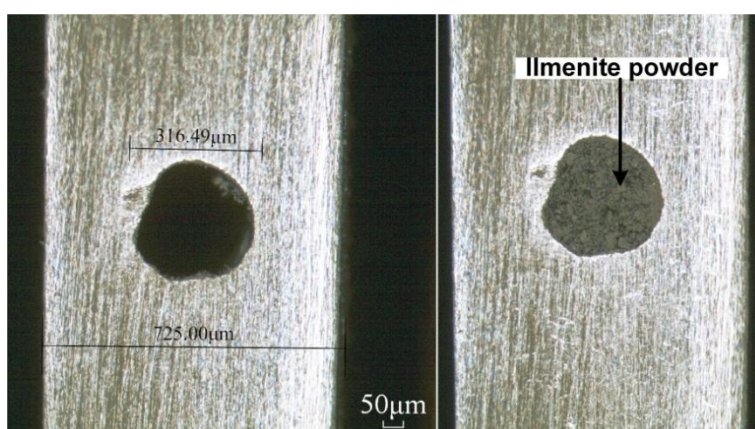


Figure 1. The optical photographs of MCE

2.3 Constant cell voltage electrolysis in the two electrodes system

In order to verify the reduction path, experiments were carried out in the two electrodes between the assembled FeTiO_3 pellet cathode and the graphite anode for different electrolysis periods

at constant cell voltage of 3.0V. The synthetic FeTiO_3 powder [49] (~2.0g) mixed thoroughly with ammonium bicarbonate (NH_4HCO_3 , 1g, served as pore-forming agent) was pressed (15MPa) in a columnar module for 5 minutes in order to form a pellet of 13mm in diameter and 3mm in thickness. The mixed powder was then sintered in a numerical controlled tube furnace in argon inert atmosphere at 1173K for 3h. Finally, the sintered pellet was wrapped with Mo wires to assemble as the cathode.

2.4 Characterization of samples

After terminations of electrolysis at prescribed periods, the cathode was removed from the cooled molten salt, and then washed with distilled water in an ultra-sonic water bath to dissolve the solidified salts. The reduced samples were then examined by X-ray diffraction (Version. D/max-2200pc) with Cu Ka radiation at a scan rate of $10^\circ/\text{min}$ in the range of $2^\theta=10\sim 90^\circ$. The morphology, particle size and elemental constituents of the reduced samples were analyzed by SEM and EDS (XL 30ESEM TMP model).

3. RESULTS AND DISCUSSION

3.1 Cyclic voltammetry

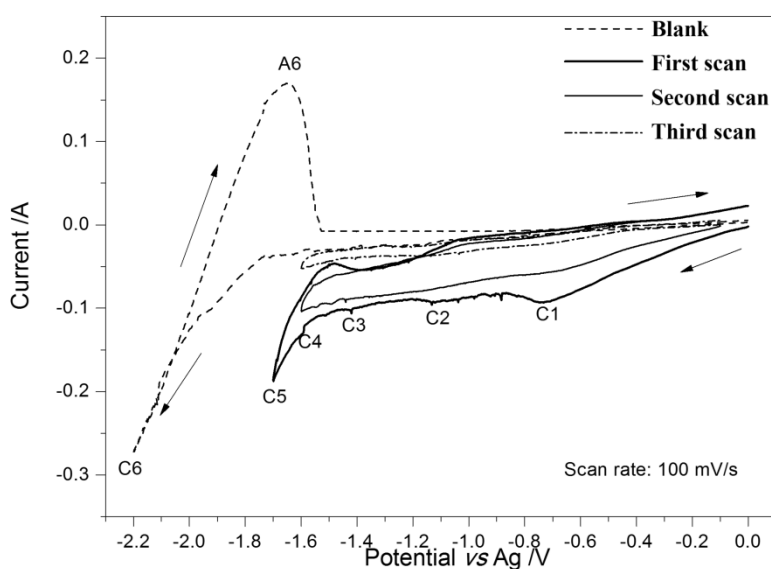


Figure 2. Cyclic voltammogram of the MCE without and with (solid line) FeTiO_3 powder of the first to third cycles in a smaller negative potential range (the first cycle: 0 to -1.7V, the second and the third cycle: 0 to -1.6V) at a scan rate of 100 mV/s.

The electro-reduction behavior of FeTiO_3 powder in the pre-electrolyzed fused $\text{CaCl}_2\text{-NaCl}$ eutectic at 973K was investigated by cyclic voltammetry measurements, as shown in Fig. 2. In comparison with the bare Mo electrode (fine dashed line) which shows the redox couples (labeled as

C6/A6) due to the deposition and re-dissolution of Ca and/or Na metal, it can be seen from Fig. 2 that the first cycle of the FeTiO_3 powder (bold solid line) presents at least four reduction peaks that is started by the decomposition of ilmenite, labeled as C1 (-0.72V vs Ag), and then followed by another three small but detected current peaks of C2 (-1.15V), C3 (-1.4V) and C4 (-1.58V) before the deposition of Ca and/or Na metal. This result indicates that the electro-reduction of FeTiO_3 is a stepwise process before the final product FeTi formed. In the second cycle, however, the current curves can be hardly found before the negative potential limits (-1.6V), indicating the irreversibility of the deoxidization of ilmenite on the continuous CVs. After the third cycle of FeTiO_3 -MCE in molten salt, the recorded currents are close to background values, suggesting the deoxidization process is almost completed. Thus, based on the above analysis, the electrochemical deoxidization of ilmenite goes through at least four steps, the intermediates are analyzed by the next electrolysis experiments.

3.2 Constant Voltage Electrolysis in a Two-Electrode Cell

In order to better understand the electrochemical reduction process of ilmenite, experiments were conducted with the constant cell voltage of 3.0V that was below the thermodynamic decomposition potentials of CaCl_2 (3.33V vs S.H.E) and NaCl (3.36V vs S.H.E) eutectic for 5min, 1h, 2h, 8h, 16h, and 20h at 973K. The intermediate products are analyzed by XRD in Fig. 3 and the inter-reactions with the theoretical electrode potentials (vs S.H.E) are listed in Table 1 also. It shows clearly that the intermediate products are Fe and CaTiO_3 mixtures reduced for 5 minutes, suggesting that FeTiO_3 particles are firstly electro-deoxidized to Fe and CaTiO_3 via reaction 1 in the initial electrolysis period. In the mixtures reduced for 1h and 2h, Fe_2Ti phase is detected. This result demonstrates that the intermediate CaTiO_3 formed is further reduced near the Fe phase via reaction 2 to obtain Fe_2Ti phase.

It is necessary to note that the initial transformation mechanism of FeTiO_3 to Fe/ CaTiO_3 as well as the following formation of Fe_2Ti through the further deoxidization of CaTiO_3 is different with the previous findings [14, 39-41]. In their studies, two steps are needed to form Fe-Ti alloys. The first step is the deoxidization of FeTiO_3 to form Fe, TiO_2 , and O^{2-} , and the spontaneous combination of them in electrolytes to form CaTiO_3 ; the second step is the deoxidization of CaTiO_3 to produce Ti metals, which will further react with Fe metals to form Fe-Ti alloys, such as Fe_2Ti and FeTi. However, the main problem is the intermediate phases of TiO_2 could be hardly detected in their XRD results and the reduction kinetics of CaTiO_3 to Ti metals is very slow. Actually, we believe that CaTiO_3 phases are formed directly with Ca^{2+} participated by reaction 1 at -0.72V vs Ag. After that, CaTiO_3 will be preferentially reduced on the basis of Fe enriched phases to form Fe_2Ti by reaction 2. This result is also evidenced by the products after 2h of electrolysis.

As furthering the electrolytic time for 8h, there's still some CaTiO_3 phases are detected. The newly formed FeTi phase is formed possibly via reaction 3 due to the more positive theoretical potential of -1.50 V (vs S.H.E), indicating the FeTi phase can be formed near Fe phase by electrochemical reduction of perovskite under the given conditions.

Table 1. The intermediates detected by XRD reduced for different time at cell voltage of 3.0V and the corresponding cathodic electrode reactions with the theoretical potentials at 973K as well as the derived potentials on the CVs.

Electrolysis time	Phases detected by XRD analysis	Possible cathodic electrode reactions	
5min	Fe, CaTiO ₃ , FeTiO ₃ , NaCl	FeTiO ₃ (s)+Ca ²⁺ +2e- =CaTiO ₃ (s)+Fe(s)	1
1h	Fe, CaTiO ₃ , Fe ₂ Ti	CaTiO ₃ (s)+2Fe(s)+4e- =Fe ₂ Ti(s)+Ca ²⁺ +3O ²⁻	2
2h	Fe, CaTiO ₃ , Fe ₂ Ti	CaTiO ₃ (s)+Fe(s)+4e- =FeTi(s)+Ca ²⁺ +3O ²⁻	3
8h	CaTiO ₃ , Fe ₂ Ti, FeTi, Ti	CaTiO ₃ (s)+4e- =Ti(s)+Ca ²⁺ +3O ²⁻	4
16h	Fe ₂ Ti, FeTi, Ti		
20h	FeTi, Fe ₂ Ti	Fe ₂ Ti(s)+Ti(s)=2FeTi(s)	5

Moreover, the intermediate CaTiO₃ can be also deoxidized to form titanium metal via reaction 4, as is detected by XRD analysis in Fig. 3d and e. When electrolysis time reaches to 16h, CaTiO₃ is found to be reduced completely by reactions 2-4 to form FeTi, Fe₂Ti and Ti. In this period, the reduced Fe₂Ti and Ti are combined chemically to form FeTi particles by reaction 5. As electrolysis for 20h, the products contain the main FeTi phase and minor Fe₂Ti phase.

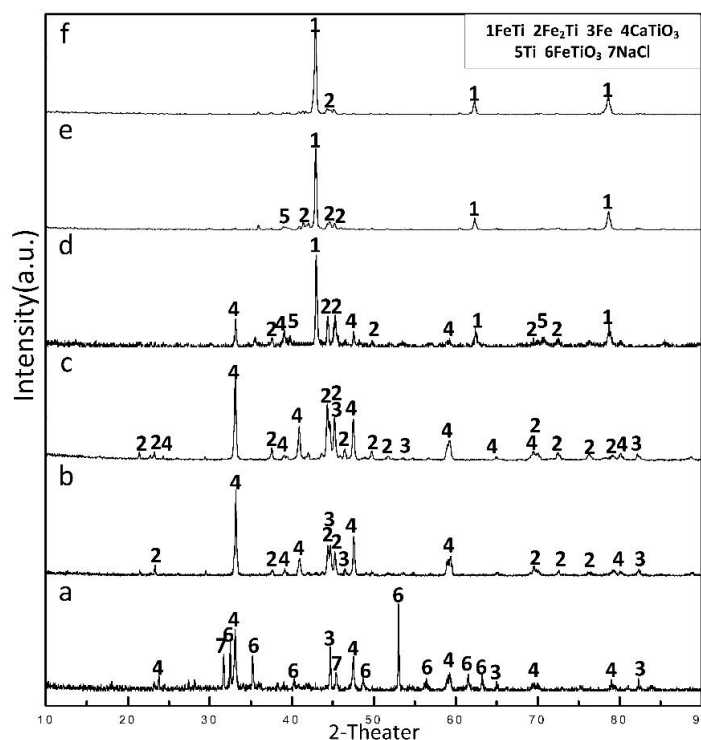


Figure 3. XRD patterns of the reduced samples at different electrolysis periods: (a) 5min, (b) 1h, (c) 2h, (d) 8h, (e) 16h, and (f) 20h.

Thus, from the XRD analysis in Fig. 3 and Table 1, it can be concluded that the electrochemical reduction of ilmenite is a stepwise process; particularly when concerning the reduction of perovskite, which is a definite intermediate product during the electrolysis of ilmenite in CaCl₂-NaCl eutectic, three main complex procedures containing Fe₂Ti, FeTi and Ti formed take place at 3.0V of electrolysis at 973K. The result is in good agreement with the electrochemical analysis in Fig. 2, that the reduction of ilmenite undergoes four intermediate procedures and thus, the recorded four peaks C1 (-0.72V vs Ag), C2 (-1.15V), C3 (-1.4V) and C4 (-1.58V) can be referred to the formation of Fe/CaTiO₃, Fe₂Ti, FeTi, and Ti, respectively.

3.3 Electron microscopy

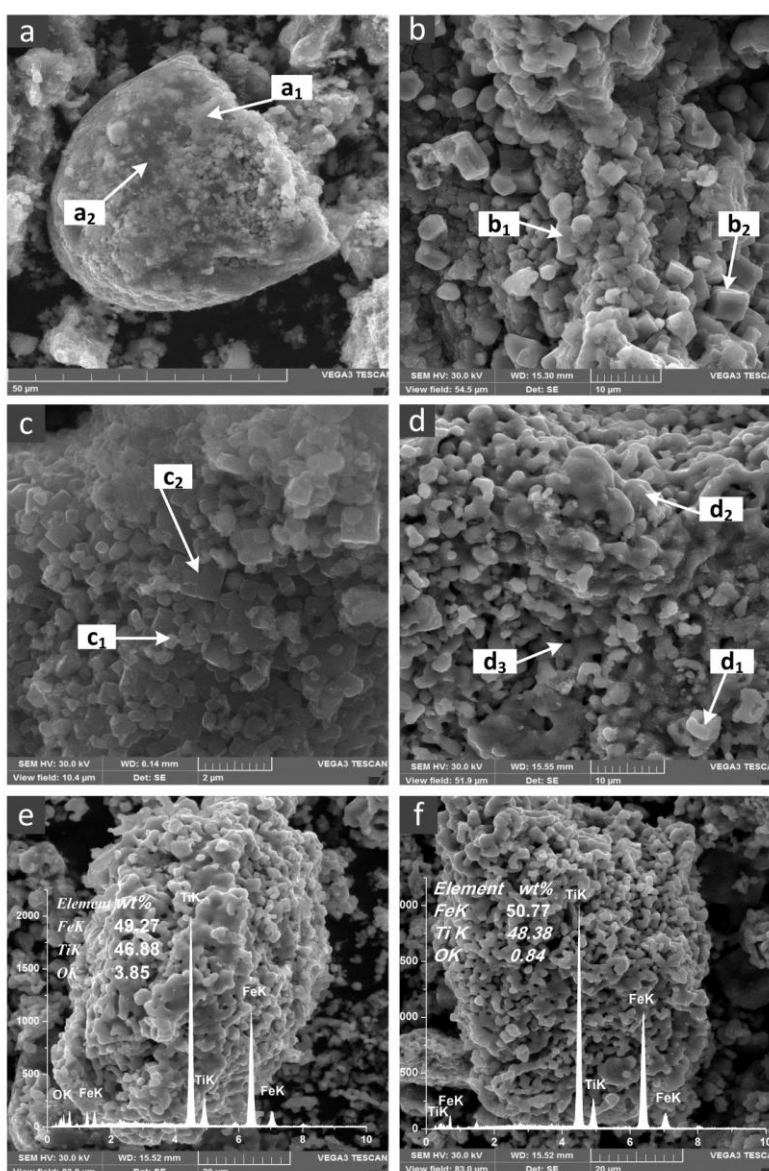


Figure 4. SEM images in partially reduced FeTiO₃ particles removed at specific intervals during the reduction process: (a) 5min, (b) 1h, (c) 2h, (d) 8h, (e) 16h, (f) 20h. The insert patterns in e and f are the corresponding EDS spectrum results measured over the image area.

The mixtures partially reduced for different electrolytic time with cell voltage of 3.0V at 973K were washed, grinded and then examined with SEM and EDS. The corresponding secondary electron images and EDS identified results of the samples are shown in Fig. 4 and Table 2. It can be seen from Fig. 4a that the new mixtures partially reduced for just 5 minutes are found to be agglomerated outside the dark gray ilmenite phase.

By EDS analysis, the composition of the new mixtures is 25.06 mass% Fe, 22.91 mass% Ti, 14.11 mass% Ca and 37.92 mass% O; and that of the dark gray ilmenite phase is 14.43 mass% Fe, 33.80 mass% Ti, 5.13 mass% Ca and 45.02 mass% O. This result demonstrates that the electrochemical reduction takes place first at the interface between the ilmenite pellet and the electrolyte. In the mixture reduced for 1h, the particle size of the reduced sample is observed to decrease to less than 10 micrometers, as seen in Fig. 4b. EDS analysis shows a mixture of 80.18 mass% of Fe and 6.6 mass% of Ti on the surface of the reduced particle, thus this mixture can be regarded as Fe₂Ti and Fe enriched phases. Table 2 lists the detailed data by EDS analysis. The 6.6 mass% of titanium embraced in the iron enriched phase implies that the perovskite is reduced to form Fe₂Ti on the basis of iron substrate via reaction 2, while the rest of perovskite phase is found to disperse with cubic crystal structures in the pellet, which is shown in Fig. 4b. As prolonging the electrolysis time for 2h, some dense structures containing 41.66 mass% Fe, 24.69 mass% Ti in fine particles and 25.49 mass% Ti, 31.88 mass% Ca, 33.42 mass% O in larger bulk are found in Fig. 4c. The dense structures, however, are not easily decomposed in spite of the large overpotentials versus Ag.

Table 2. XRD phases detected at various electrolysis periods and the corresponding elemental compositions from Fig. 2.

Sample number	Electrolysis time	Detected area	EDS analysis			
			Fe mass%	Ti mass%	Ca mass%	O mass%
a	5min	a ₁	25.06	22.91	14.11	37.92
		a ₂	14.43	33.80	5.13	45.02
b	1h	b ₁	80.18	6.60	8.44	3.41
		b ₂	20.20	26.6	28.44	23.41
c	2h	c ₁	41.66	24.69	15.49	8.78
		c ₂	9.91	25.49	31.88	33.42
		d ₁	42.44	41.94	9.95	9.33
d	8h	d ₂	36.57	24.19	15.49	12.41
		d ₃	1.37	40.70	33.37	13.76
		full area	49.27	46.88	0.2	3.85
f	20h	full area	50.77	48.38	<i>undetected</i>	0.84

According the XRD analysis in Fig. 3, the decomposition of CaTiO₃ are occurred in the presence of Fe phases, indicating the pre-formed Fe actually decrease the decomposition of CaTiO₃, as a result, Fe₂Ti and FeTi are directly produced. After electrolyzed for 8h, more ferrotitanium alloy grains are produced and some pores of about 3 micrometers are also present in the samples. The metallic grains increase with the prolonging time in the electrolytic durations. At 16h of reduction, the porous structures are formed, as shown in Fig. 4e. The EDS analysis determines the composition of the

porous particles to be FeTi with 3.43 mass% O. Comparing the oxygen content after reduced for 8h and 16h, the furthering 8h of electrolysis results in the removal of very little oxygen content. This demonstrates the slow reduction kinetics during the deoxidization process, that is, the diffusion of O^{2-} would be a limiting step in the reduction process. When reduction time reaches to 20h, the SEM image in Fig. 4f shows a porous structure of FeTi with 0.17 mass% O in 20 micrometers particle size, indicating a complete reduction of ilmenite via the heterogeneous reactions 2-5 by electrochemical deoxidization method.

4. CONCLUSIONS

(1) Ferrotitanium alloys have been successfully prepared from ilmenite by the electrochemical reduction process, using equal-molar $CaCl_2$ -NaCl metls as electrolyte, ilmenite pellet tied by molybdenum rod as cathode and graphite as anode at 973K with cell voltage of 3.0V that is below the decomposition of molten salt.

(2) Cyclic voltammetry and constant cell voltage electrolysis in the two electrodes system together with XRD, EDS analysis and SEM observation suggest that the electrochemical reduction of ilmenite is a stepwise process rather than a single one. During the reduction, $FeTiO_3$ is first electrochemically reduced to Fe/ $CaTiO_3$ mixtures. The intermediate $CaTiO_3$ is further reduced preferentially in the presence of solid Fe to form ferrotitanium alloys (Fe_2Ti and FeTi) as well as directly deoxidized to titanium metal, which is then combined with Fe_2Ti to form ferrotitanium.

(3) The micromorphology evolutions of the reduced samples have been addressed. From the SEM observation, it can be seen that the electrochemical reduction takes place first at the interface between ilmenite and the electrolyte. As the electrolysis process inward, fine but dense structures containing Fe_2Ti and $CaTiO_3$ mixtures have been found. The dense particles could block the diffusion path of O^{2-} and eventually decrease the deoxidization rate. Thus, the pores presented in the reduced particles play an important role for the transport of O^{2-} in the deoxidization process.

(4) The removal value of oxygen shows a slow deoxidization rate in the whole process. This demonstrates that the reduction kinetics is strongly depended on the O^{2-} diffusion rate.

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